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Algebra 2 P Chapter Practice Test Name: Please show your work. You are allowed only a basic calculator. This is last years test.1. Perform the indicated operations and simplify 2a) $(5x^2 - 7y)$ b) $(2x + 7)(3x^2 - 8x + 1)$ 2. Factor each expression completely a) $x^2 - 3x - 40$ b) $2x^3 + x^2 - 6x - 3$ c) $x^4 + 5x^3 + 6x^2$ d) $2(a + b)^2 + 5(a + b) - 33$. a) Write $?x < 8$ in interval notation and...

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[UNIT 4 PRACTICE TEST](#)

Honour Chemistry: Unit 4 Practice Test: Thermochemistry and Nuclear ChemistryPart A: Multiple Choice and Numerical Response (1 mark each for Multiple Choice) (2 marks each for Numerical Response) $E_{\text{bind}} = ?mc^2$ $c = \text{speed of light}$ (3.00 10⁸ m s⁻¹) $1 \text{ eV} = 1.69 \cdot 10^{-19} \text{ J}$ (1 kg = 6.022 10²⁶ amu , or 1 g = 6.022 10²³ amu = 1 mole amu) $1 \text{ MeV} = 1.69 \cdot 10^{-16} \text{ J}$ (1 Neutron = 1.008665 amu) (1 Proton = 1.007825 amu)

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[UNIT 2 \(PART 1\) PRACTICE TEST](#)

Honour Chemistry Practice Test: Unit 2 (Part 1): Matter as Solutions and Gases $1 \text{ atm} = 760 \text{ mm Hg} = 760 \text{ torr} = 101.325 \text{ kPa}$ $TK = TC + 273.15$ $STP = 22.4 \text{ L mol}^{-1}$ @ 0.00C & 101.325 kPa (1 atm) $SATP = 24.8 \text{ L mol}^{-1}$ @ 25.0C & 100. kPa $L \text{ atm} = L \text{ kPa} \cdot m$ $n_{\text{solute}} R = 0.0821$ or 8.314 $D = n = C = K \text{ mol} \cdot K \text{ mol} \cdot V \cdot M \cdot V_{\text{solution}}$ $P_1 V_1 = P_2 V_2$ $r_1 M_2 = n_1 P_1 C_1 V_1 = C_2 V_2$ $PV = nRT = = ?1 = = n_1 T_1 = n_2 T_2 = r_2 M_1 = n_{\text{Total}} P_{\text{Total}} P_{\text{Total}} = P_1 + \dots$

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[PRACTICE TEST 7](#)

Practice Test 7, solutions belowProblem 11. (p & q) A2. (p v q) A3. Show: rProblem 21. (p ? q) ? (r ? s) A2. p v q A3 (q & p) A4. Show: (r v s)Problem 31. (p v q) v (r v s) A2. r A3. (t ? s) A4. Show p ?q Problem 11. (p & q) A2. (p v q) A3. Show: r4. p & q &, 25. p & q , 46. p & q , 57. r !, 1,6Problem 21. (p ? q) ? (r ? s) A2. p v q A3 (q & p) A4. Show: (r v s)5. (r v s) A1P6. p ? q ?v, 27. q v ...

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